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MUST be able to write a formula without a second thought.

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Mass of proton =  $1.6726 \times 10^{-24}$  g

Mass Page 5/9. Access Free Amount



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Of Substance Chemsheets of electron  
=  $9.1094 \times 10^{-28}$  g Mass of neutron =  
 $1.6749 \times 10^{-24}$  g Avogadro constant  
=  $6.022 \times 10^{23}$  mol<sup>-1</sup>. a) Calculate  
the mass of a 1H atom. b)

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5.266 moles of  $\text{TiCl}_4$  needs 10.53 moles of Mg to react,  $\text{TiCl}_4$  is in XS and does not all react, so Mg is the limiting reagent 2.058 moles of  $\text{TiCl}_4$  reacts with 4.115 moles of Mg 2.058 moles of Ti is produced

Mass of Ti =  $2.058 \times 47.9 = 98.6 \text{ g}$ . ©

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a i M r MgO = 40.3 0.741/40.3 =  
0.0184 1 1 If used 40 then penalise  
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amount = mass/Mr = 5 / (23.1 x2 + 12 +16 x3) = 0.0472 mol conc=

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amount/Volume =  $0.0472 / 0.25 = 0.189 \text{ mol dm}^{-3}$

Example 1: What is the amount, in mol, in 35.0g of  $\text{CuSO}_4$ ? amount = mass/Mr =  $35 / (63.5 + 32 + 16 \times 4) = 0.219 \text{ mol}$  It is usually best to give your answers to 3sf

Example 3 : What is the volume in  $\text{dm}^3$  at room

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5. Formulae, equations and amounts of substance

Welcome to Topic 2 - AMOUNT OF SUBSTANCE. Topic 2 specification content Topic 2 notes

Topic 2 - Amount of Substance - A-

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## Level Chemistry

In chemistry, the amount of substance in a given sample of matter is defined as the number of discrete atomic-scale particles in it divided by the Avogadro constant  $N_A$ . In a truly atomistic view, the amount of substance is simply the number of particles that constitute the



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substance. The particles or entities may be molecules, atoms, ions, electrons, or other, depending on the context. The value of the Avogadro constant  $N_A$  has been defined as  $6.02214076 \times 10^{23} \text{ mol}^{-1}$ . In the truly atomistic ...

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