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Chapter 6 Answers Page 195 #1. Use the tables on page 420 (heat of dissolution) and 422 (polyatomic ions) to help you answer these questions. a) $\text{HClO}_3(\text{s}) \rightarrow \text{H}^+(\text{aq}) + \text{ClO}_3^-(\text{aq})$ $\Delta H =$ b) $\text{HI}(\text{s}) \rightarrow \text{H}^+(\text{aq}) + \text{I}^-(\text{aq})$ $\Delta H =$ c) $\text{KNO}_3(\text{s}) \rightarrow \text{K}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$... <https://studylib.net/doc/5610530/chemistry-chapter-6-answer-key>.

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Modern Chemistry 47 Chemical Bonding CHAPTER 6 REVIEW Chemical Bonding SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. _____ In metals, the valence electrons are considered to be (a) attached to particular positive ions.

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CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

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SECTION 6.1 Nature favors arrangements in which potential energy is minimized. For example, a boulder is less likely to balance at the top of a hill than it is to roll to the bottom of a valley. The boulder at the top of the hill is not stable. Atoms are usually not stable when isolated. They usually combine to form more stable arrangements of matter.

[CHAPTER 6 Chemical Bonding](#)

molar mass of $\text{C}_6\text{H}_6\text{O} = 94.11 \text{ g}$ 21. a. molar mass of $\text{SO}_3 = 80.07 \text{ g}$ 49.2 mg SO_3 1 g 1000 mg 1 mol 80.07 g = 6.14 10^{-4} mol SO_3 3 b. molar mass of $\text{PbO}_2 = 239.2 \text{ g}$ 7.44 $\times 10^4 \text{ kg PbO}_2$ 1000 g 1 kg 1 mol 239.2 g =

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3.11 105 mol PbO 2 c. molar mass of $\text{CHCl}_3 = 119.37 \text{ g}$ 59.1 g CHCl_3 1 mol 119.37 g = 0.495 mol CHCl_3 d. molar mass of $\text{C}_2\text{H}_3\text{Cl}_3 = 133.39 \text{ g}$ 3.27 ...

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